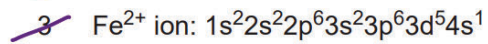
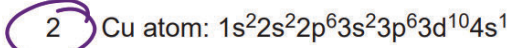
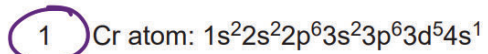


1. Which electron configuration(s) is/are correct?



more stable to partially fill 4s subshell

A 1, 2 and 3

B Only 1 and 2

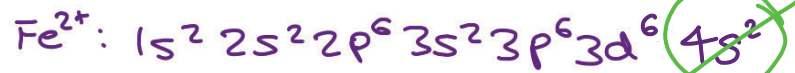
C Only 2 and 3

D Only 1

Your answer

B

ions do not follow the same pattern:



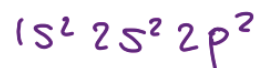
loses 4s electrons first [1]

2. In the diagrams below, each box represents an orbital and each electron is shown as an arrow.

Which diagram shows the correct arrangement of electrons in an atom of carbon?

A	
B	
C	
D	

electron configuration
of carbon:



1s 2s 2p

more stable to
have 2 partially
filled subshells
(less electron repulsion)

Your answer

D

[1]